

# Chemical limnology

Limnology

Lecture 8

# Outline

- Carbon cycle
- Nutrients
- Toxic chemicals

# pH

Measure of hydrogen ion concentration

Expressed as  $10^x$  moles of hydrogen ions per liter

$\text{pH} = -\log_{10}(\text{H}^+ \text{ ions})$

$0.0000001 = 1 \times 10^{-7}$  moles  $\text{H}^+$

$-\log_{10}(1 \times 10^{-7})$

$\text{pH} = 7$

$\text{pH} < 7$  acid

$\text{pH} = 7$  neutral ( $\text{H}^+$  and  $\text{OH}^-$  equal)

$\text{pH} > 7$  basic

# Sources and forms of carbon in Lakes

## Sources

Atmosphere

Rocks (limestone and dolomite) and soils

Detritus – dead organic material

## Forms

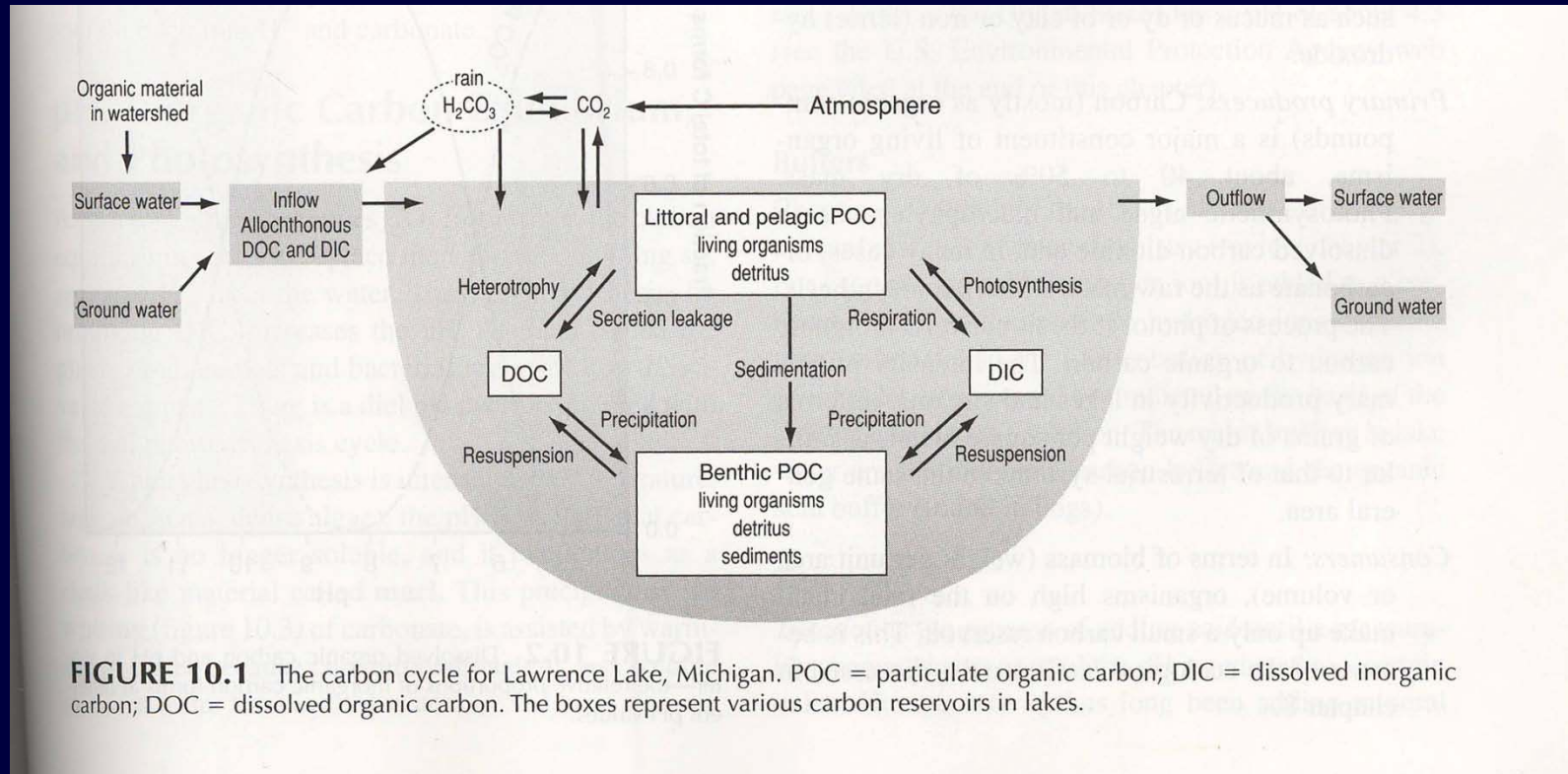
Dissolved Inorganic Carbon (DIC) –  $\text{CO}_2$

Dissolved Organic Carbon (DOC) – methane, humic acids

Particulate Organic Carbon (POC) – living, dead organisms, feces

# Carbon cycle

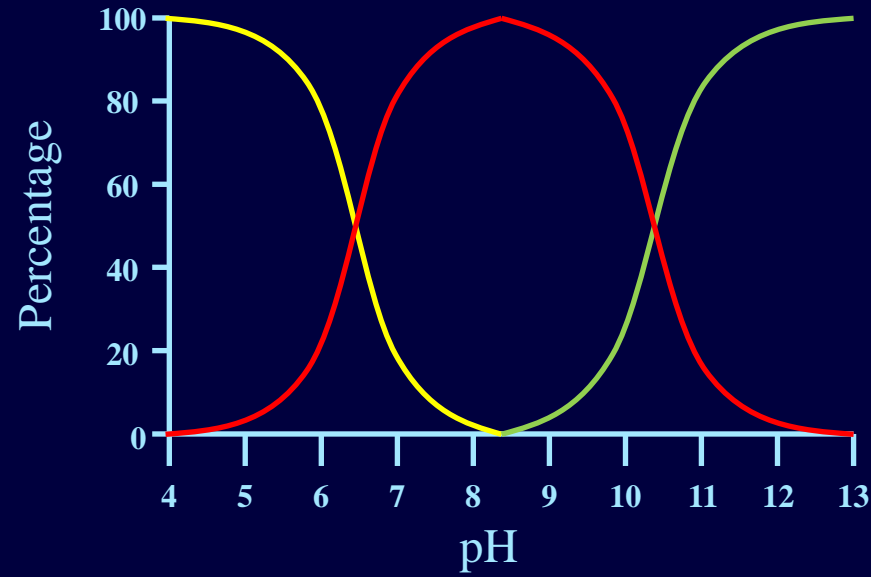
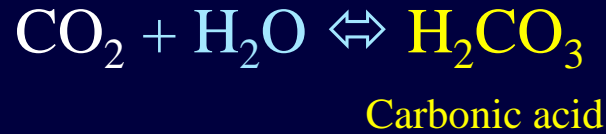
- Fig. 10.1



# Carbon Dioxide in Lakes

- Critical: Photosynthesis
- Highly soluble –  $\text{CO}_2$  200x more soluble than  $\text{O}_2$

# Carbon Dioxide in Lakes

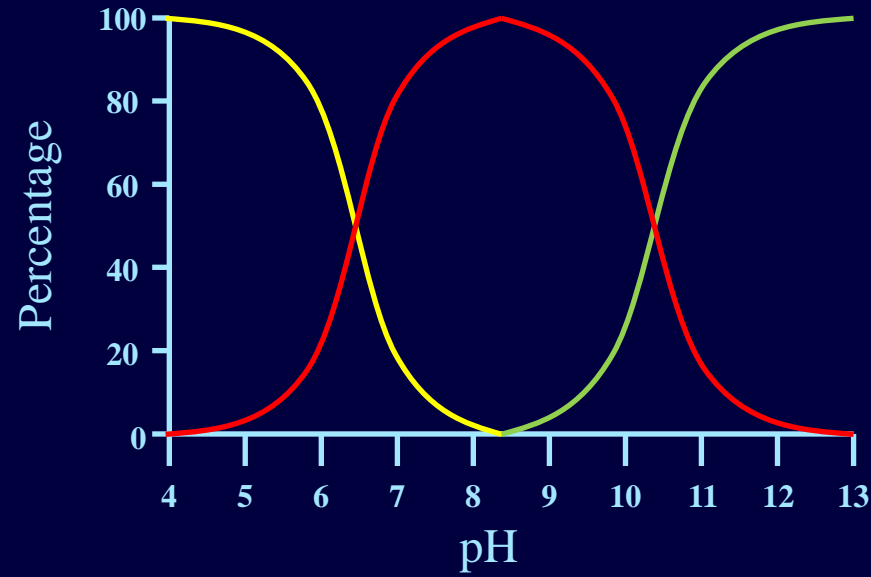


Most lakes

# Carbon Dioxide in Lakes



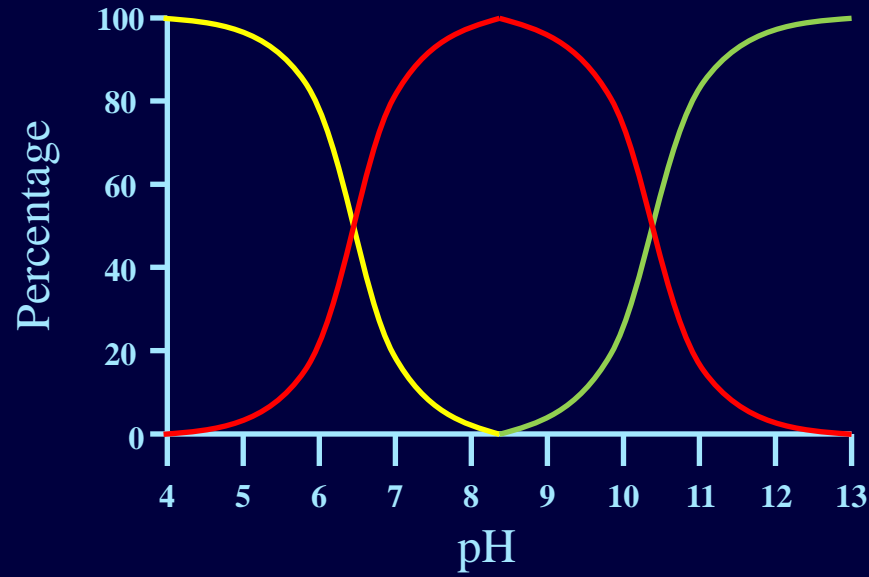
Carbonic acid    Bicarbonate



Most lakes



# Carbon Dioxide in Lakes



Most lakes

# Carbon Dioxide in Lakes



Add CO<sub>2</sub>, increase H<sup>+</sup>, decrease pH

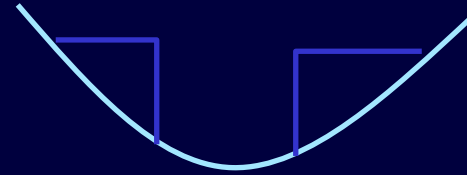
Remove CO<sub>2</sub>, decrease H<sup>+</sup>, increase pH

# Marl Lake (also known as Bench Lake)

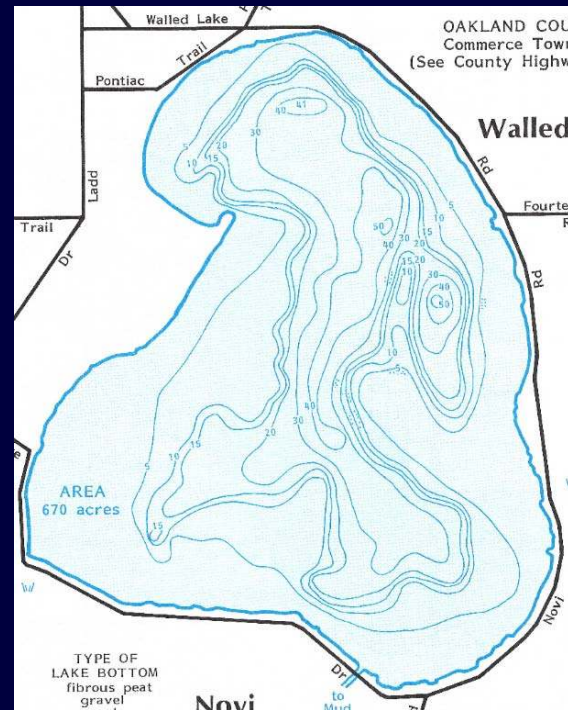


Calcium bicarbonate  
(highly soluble)

Calcium carbonate  
(low solubility)



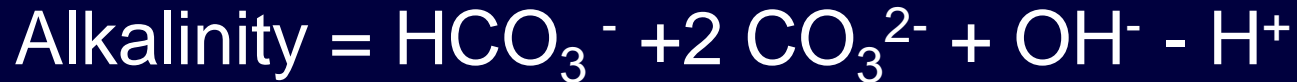
'Marl Lakes' or 'Bench Lakes'



Shallow margin created by Calcium Carbonate deposition over thousands of years

# Carbon Dioxide in Lakes

Alkalinity – buffering capacity of water relative to acid input



# Limiting nutrients

Nitrogen (amino acids)

Phosphorus (ATP)

# Nitrogen

## Forms

Nitrogen gas ( $N_2$ )

inert, not bio-available

Nitrate ( $NO_3^-$ )

bioavailable

Nitrite ( $NO_2^-$ )

toxic

Ammonia ( $NH_3$  in water  $NH_4^+$ )

Nitrogen fixation from bacteria

Excretion

## Sources of N

Fixation by bacteria (e.g., cyanobacteria)

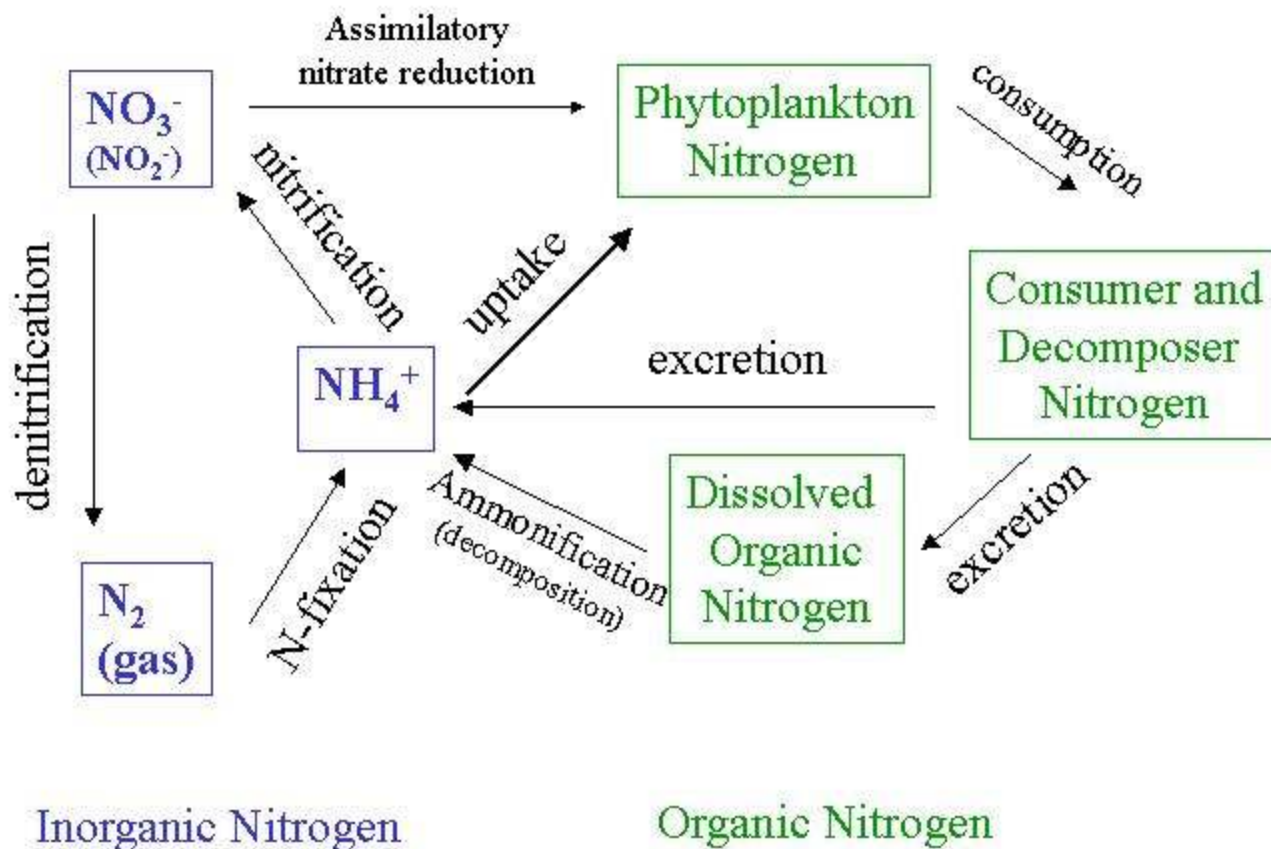
Lightning/precipitation

Weathering of minerals/runoff

Anthropogenic N

## Denitrification

# Aquatic Nitrogen Cycle



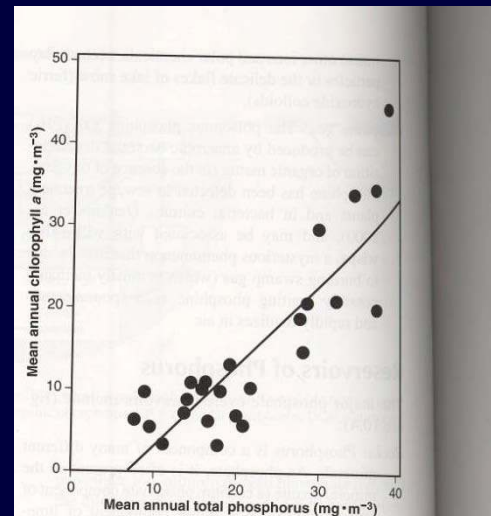
# Phosphorous

## Sources of P

- Weathering of minerals (deep ocean, coastal sediments)
- Lake sediment
- Decomposition and excretion (major source)
- Anthropogenic P greater than natural inputs of P



Experimental Lakes Area, Lake 227  
Experimental addition of P



**FIGURE 10.5** The relationship between average annual concentrations of total phosphorus and chlorophyll a (a measure of algal abundance) in lakes of the Experimental Lakes Area, Ontario, Canada. See figure 8.4 for the relationship between fish production and phosphorus concentration. The wide range of phosphorus concentrations in this figure was made possible by including data from fertilized and unmodified lakes. The correlation explains about 74% of the variation in algal abundance. Source: data from Schindler, 1977.



# Nutrients

Phosphorus (ATP)

Nitrogen (amino acids)

Redfield ratio: normal ratio in water

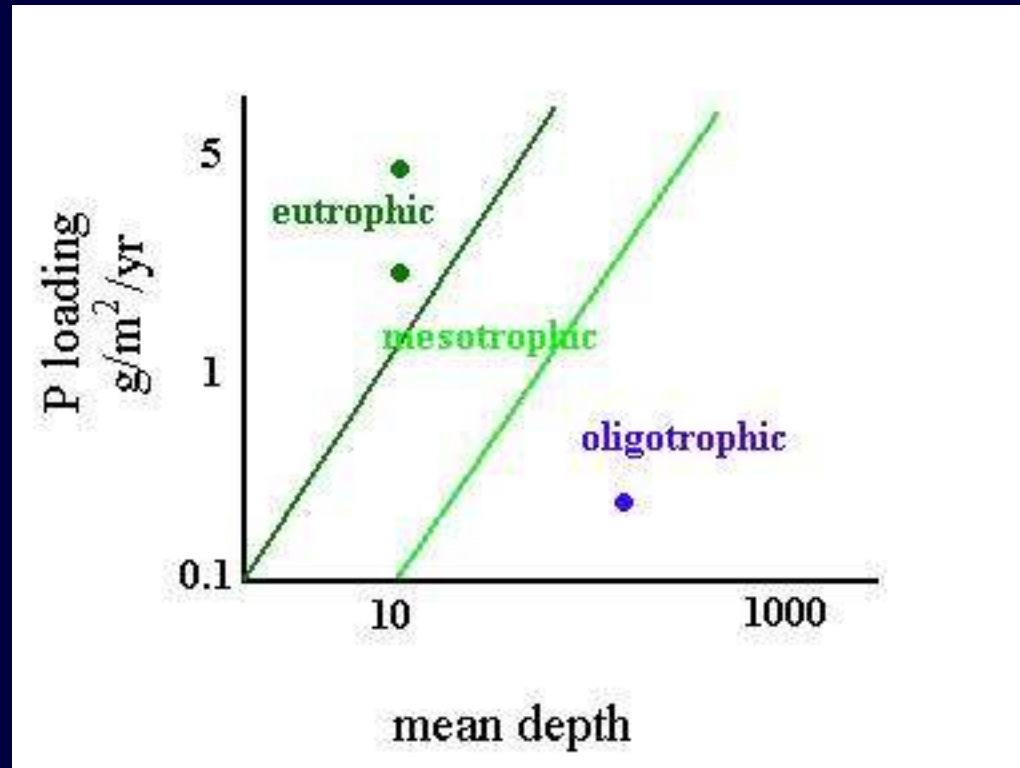
oceans → 106 C : 16 N : 1 P

lakes → 400 C : 30 N : 1 P

Phosphorous often limiting (i.e., available in less quantities than required by organisms)

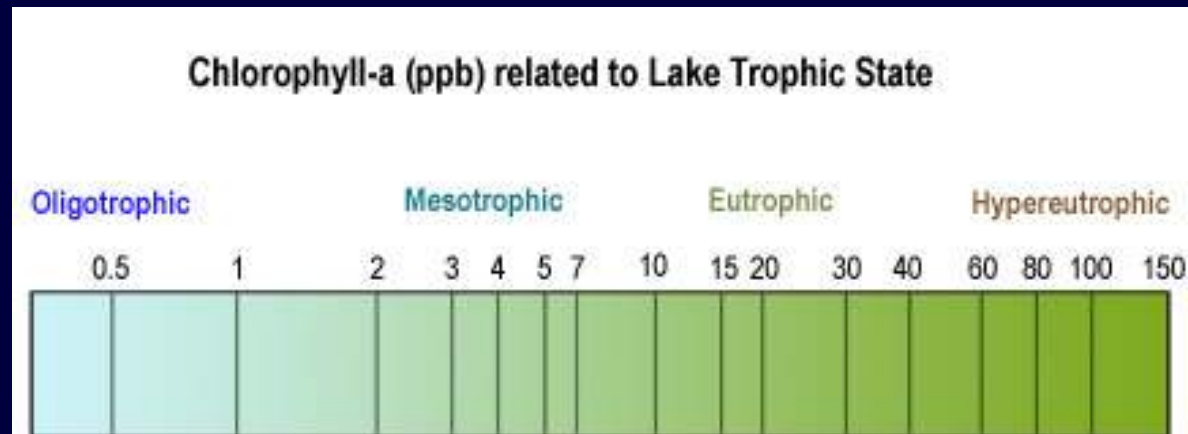
? What does this mean practically?

# Phosphorus and productivity



K. Shulz

# Lake trophic status



Oligotrophic – low nutrients, high oxygen, low productivity, clear

Eutrophic – high nutrients, anoxic hypolimnion in summer, high productivity, opaque

→ wet meadow (natural)

cultural eutrophication if human

Dystrophic – stained with DOC, low nutrients, low productivity (often small, mtn. lakes in forests)



Oligotrophic – Lake Tahoe



Eutrophic – Lake Mendota



dystrophic – Adirondacks lakes

# Lake trophic status

PHOSPHORUS AND CHLOROPHYLL CONCENTRATIONS AND SECCHI DISK DEPTHS CHARACTERISTIC OF THE TROPHIC CLASSIFICATION OF LAKES			
	Oligotrophic	Mesotrophic	Eutrophic
Total Phosphorus (mg/m <sup>3</sup> ) Average	8	26.7	84.4
Chlorophyll <i>a</i> (mg/m <sup>3</sup> ) Average	1.7	4.7	14.3
Secchi Disk Depth (m) Average	9.9	4.2	2.45

*Data from Wetzel, 2001 Table 13-18*

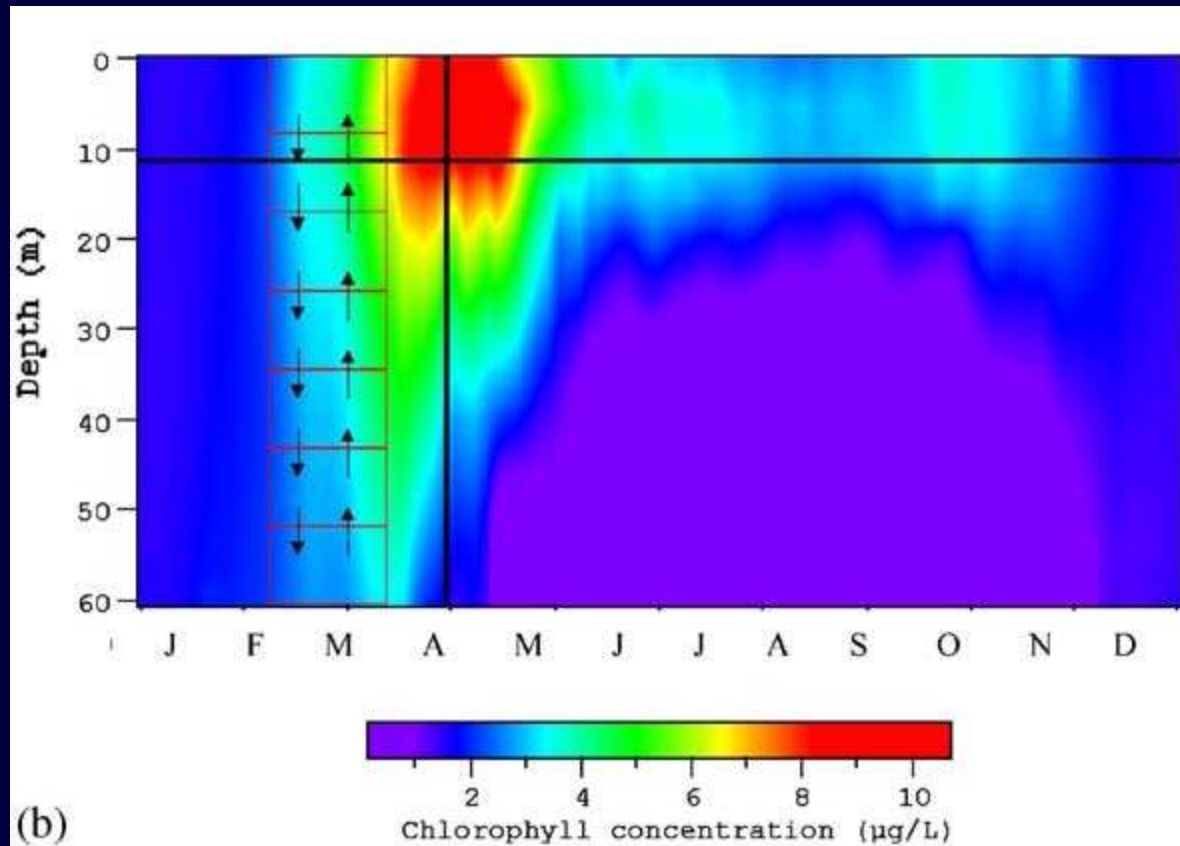
Often spring and late fall algal blooms in temperate lakes

Why?





# Mixing effects



Lake Washington, Arhonditsis et al. 2004

# Ecological stoichiometry

How ratio of different chemical elements affects ecology and vice versa

Certain species favored under certain conditions  
When N/P is low, favors cyanobacteria

Carbon-rich phytoplankton junk food of the lake  
(high calorie, no P/N)  
couch potato zooplankton can't keep up  
with phytoplankton

Shift from *Daphnia* to *Bosmina* with low P





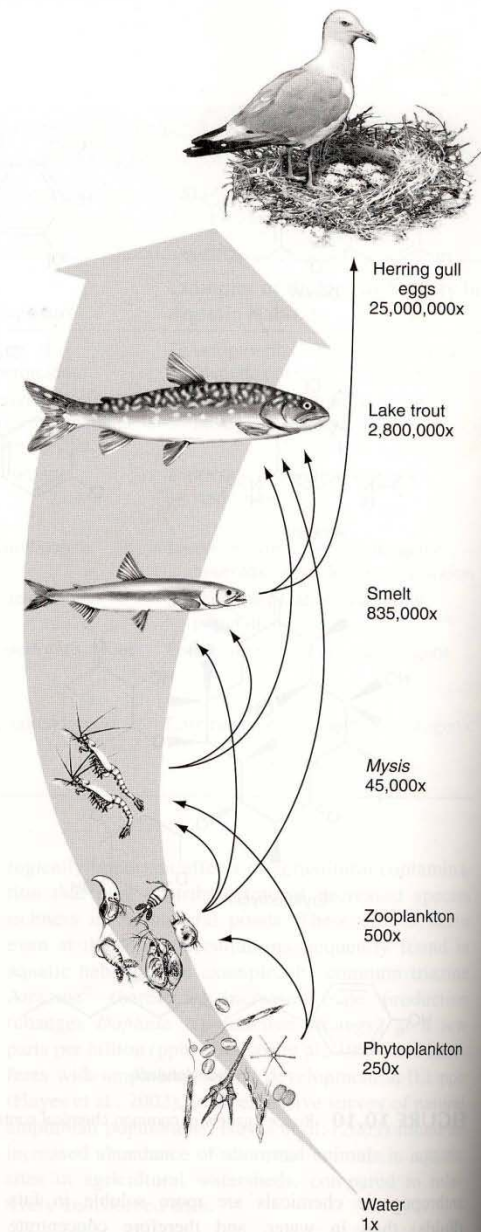
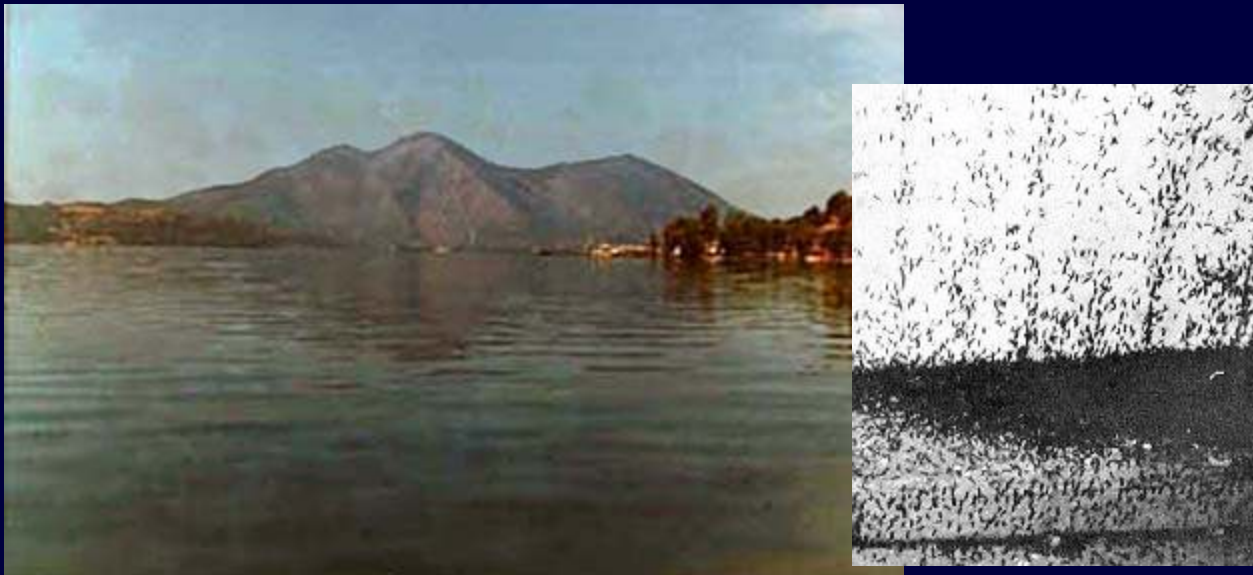
# Toxic chemicals

## Biomagnification

lipophilic chemicals concentrated  
up food chain

PCB, DDD/T, Mercury

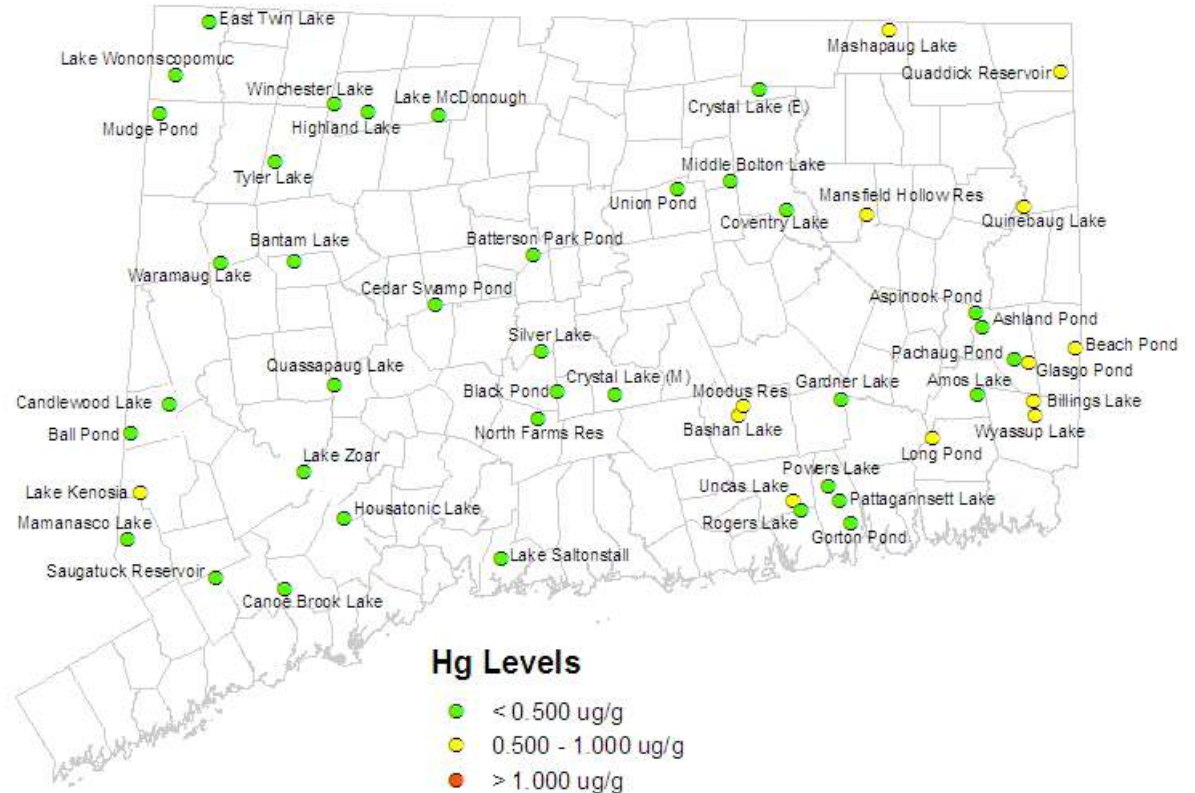
Clear Lake example



**FIGURE 10.12** An example of bioaccumulation. The concentration factors are for PCBs relative to the concentration in the water, which was estimated to have been about  $5 \times 10^{-3} \text{ g l}^{-1}$  or 5 parts per trillion (Norstrom et al., 1978). The factor of 25 million biomagnification refers to herring gull eggs, which contained about  $125 \text{ mg PCB kg}^{-1}$  wet weight. See figure 10.11 for the chemical structure of a typical polychlorinated biphenyl.

# Toxic chemicals

Mercury in CT  
most fish with Hg  
statewide consumption advisory



# Deformities

Amphibian deformities reported in many ponds in 90s

Parasitic infections in limb nodes

Nutrients increase snails → increases infection by trematode

Case closed? But high deformities in VT and no trematodes



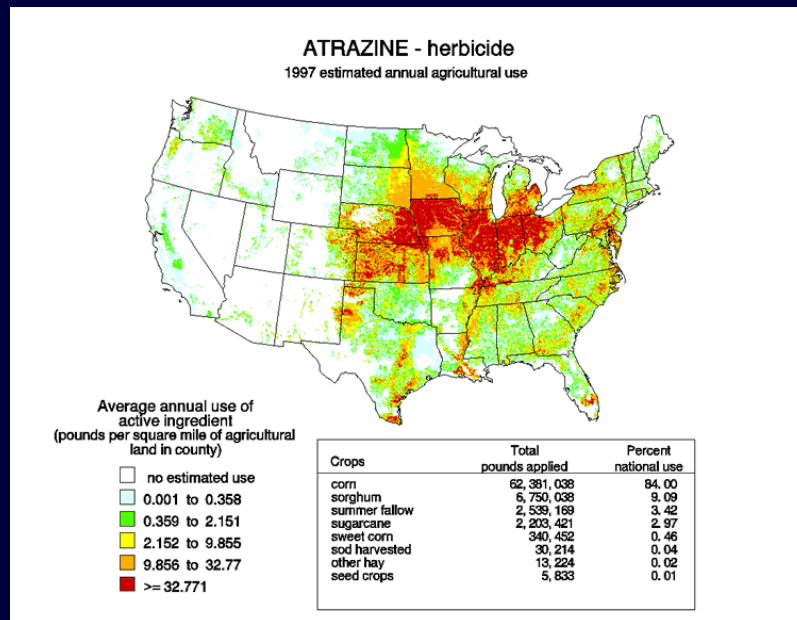
# Deformities

Amphibian males grow female ova producing cells

Atrazine (common herbicide)

Case closed? EPA panel reviews experiments and finds faults

Results not repeated . . .



Sex and suburban frogs

